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chapter outline we have examined the basic ideas of bonding showing that atoms share electrons to form molecules with stable lewis structures and that we can predict the shapes of those molecules by valence shell electron pair repulsion vsepr theory

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endothermic reaction a chemical reaction or process in which a greater amount of energy is required to break the existing bonds in the reactants than is released when the new bonds form in the product molecules heat in oxyacid any acid that contains hydrogen and an oxyanion structural formula

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chemical bonds in a perfect diamond crystal each c atom makes four connections bonds to four other c atoms in a three dimensional matrix four is the greatest number of bonds that is commonly made by atoms so c atoms maximize their interactions with other atoms

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8 1 valence bond theory valence bond theory describes bonding as a consequence of the overlap of two separate atomic orbitals on different atoms that creates a region with one pair of electrons shared between the two atoms

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skills to develop define three common types of chemical reactions precipitation acid base and oxidation reduction classify chemical reactions as one of these three types given appropriate descriptions or chemical equations identify common acids and bases

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8 7 the covalent chemical bond a model a strengths of the bond model 1 associates quantities of energy with the formation of bonds between elements 2 allows the drawing of structures showing the spatial relationship between atoms in a molecule 3 provides a visual tool to understanding chemical structure b weaknesses of the bond model 1

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as a pictorial understanding of where the electrons are in an atom we represent the electrons as dots around the symbol for the element the number of valence electrons available for bonding are

indicated by unpaired dots these symbols are called lewis symbols or lewis electron dot symbols

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what is a combustion reaction and what is its formula substance combines with oxygen releasing a large amount of energy in the form of heat and light general equations fuel o₂ co₂ h₂o

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with the background needed to understand organic reactions now covered this chapter has begun the systematic description of major functional groups alkenes are generally prepared by an elimination reaction such as dehydrohalogenation the elimination of hx from an alkyl halide or dehydration the elimination of water from an alcohol

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this video explains the concepts from your packet on chapter 8 basic concepts of chemical bonding which can be found here [goo gl tyuj36](#) section 8 1 lewis symbols and the octet

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